

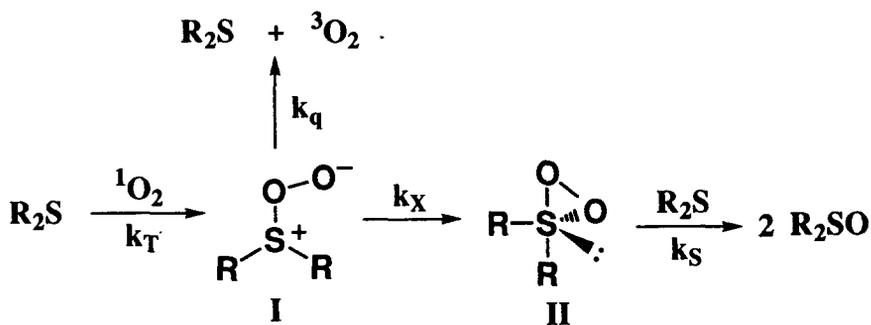
## Steric and Electronic Effects on the Partitioning of a Persulfinyl Intermediate Along the Physical Quenching and Chemical Reaction Channels.

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**Abstract:** The rate constants,  $k_T$ ,  $k_r$ , and  $k_Q$ , corresponding to the substrate induced total removal, chemical removal, and physical removal of singlet oxygen, respectively, have been measured for a series of sulfenamides. The structural effects on the partitioning of the persulfinamide intermediate along the chemical and physical deactivation pathways is discussed. Copyright © 1996 Elsevier Science Ltd

Footo and coworkers pointed out more than 25 years ago that sulfides remove singlet oxygen from solution at a significantly faster rate,  $k_T$ , than products are produced,  $k_r$ .<sup>1</sup> Furthermore, the quantum yield for the formation of diethylsulfoxide, which could theoretically be as large as 2, is in fact less than 0.05. (Scheme 1) Footo and coworkers demonstrated that it was the decomposition of a persulfoxide intermediate, I in Scheme 1, to sulfide and triplet oxygen in a physical quenching process,  $k_Q$ , which was responsible for the reaction inefficiency.<sup>2</sup>



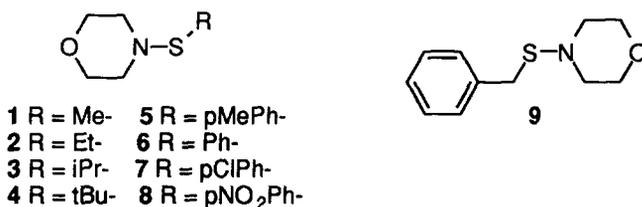
Scheme 1

With the advent of a technique to directly monitor the disappearance of  ${}^1\text{O}_2$  the values of  $k_T$ , the rate constants for substrate induced removal of  ${}^1\text{O}_2$  are now available with a high degree of accuracy.<sup>3</sup> The rate constants for product formation,  $k_r$ , are also available, often via a competition technique, with a reasonable degree of accuracy. Consequently, the remaining bimolecular rate constant of interest,  $k_Q$ , can be obtained by

difference using equation 1. The variable  $\alpha$  in equation 1 is a function of the product composition. When sulfoxide is the exclusive product  $\alpha = 0.5$  since for each molecule of singlet oxygen that disappears two molecules of product are formed. On the other hand, when sulfone is the exclusive product  $\alpha = 1.0$ . Unfortunately, as pointed out above,  $k_f$  for sulfides is very small in comparison to  $k_T$  and the difference ( $k_Q = k_T - \alpha k_f$ ) cannot be determined accurately.

$$k_T = \alpha k_f + k_Q \quad 1$$

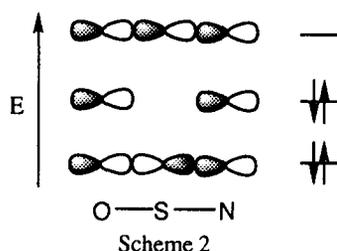
A different situation is encountered during the photooxidations of sulfenamides, **1-8**.<sup>5,6</sup> (Eqn. 2) These reactions are considerably more efficient than sulfide photooxidations and as a result  $k_T$  and  $k_f$  are of the same order of magnitude. In addition, detailed kinetic studies with 4-morpholinyl benzyl sulfide, **9**, have demonstrated that the reaction surface describing the interaction of singlet oxygen with sulfenamides is topographically similar to the Foote reaction surface for diethylsulfide (Scheme 1). Consequently, these substrates provide the first opportunity to accurately measure the steric and electronic effects on the  $k_x/k_q$  (Scheme 1) ratio.



The rate constants  $k_T$  were determined in oxygen saturated benzene by monitoring the ability of **1-8** to quench the phosphorescence of <sup>1</sup>O<sub>2</sub> at 1270 nm using the laser based emission apparatus previously described.<sup>7,8</sup> The rate constants for product formation,  $k_f$ , were determined by competition with Δ9(10)-octalin and cyclohexene for sulfenamides **1-7** and **8**, respectively, using the method of Higgins and Foote.<sup>9</sup> The rate constants for physical quenching,  $k_Q$ , were determined using equation 1. The observed rate constants,  $k_f$  and  $k_Q$ , are related to the microscopic rate constants depicted in Scheme 1 as shown in equations 2 and 3. These equations were derived from steady state treatment of both intermediates **I** and **II** in Scheme 1 and demonstrate that 1/2 of the  $k_f/k_Q$  ratio is equal the microscopic rate constant ratio  $k_x/k_q$ . This ratio,  $k_T$ ,  $k_f$ ,  $k_Q$ , and the % of the persulfenamide intermediate (**I** in Scheme 1) which partitions along the chemical reaction channel are given in Table 1 for photooxidations of sulfenamides **1-8**.



Electronic effects, on the other hand, play a substantial role in determining the overall outcome of the reaction. As the substituent on the sulfenamide is changed from pMePh- to pNO<sub>2</sub>Ph- the percent of persulfenamide which is converted to product decreases from 27 to 13%. (%  $k_x$  in Table 1) Clearly the substituent effects demand that the amount of positive charge on sulfur that disappears is greater in the transition state for physical quenching than in the transition state for conversion to the thiadioxirane. This is consistent with the molecular orbital picture of sulfuranes (Scheme 2) which places excess electron density on the terminal atoms of the 4-electron 3-centered bond and a deficiency of electron density on sulfur.<sup>12</sup>



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